Name_____

Calculate the molarity of the following solutions.

1)   2.2 moles of $LiNO_3$ are dissolved to make 750 mL of solution.

2)   0.760 moles of $C_{12}H_{22}O_{11}$ are dissolved to make 2 L of solution.

3)   4.4 moles of silver nitrate are dissolved to make 0.750 L of solution.

4)   18 moles of $Cu(NO_3)_2$ are dissolved to make 1.250 mL of solution.

5)   0.630 moles of Li_2SO_4 are dissolved to make 250 mL of solution.

6)   6.1 moles of hydrogen chloride are dissolved to make 1.25 L of solution.

7)   0.270 moles of potassium sulfate are dissolved to make 1.50 L of solution.

8)   30 moles of sodium hydroxide is dissolved to make 1.750 mL of solution.

9)   0.048 moles of sodium carbonate are dissolved to make 0.25 L of solution.

10) 4.0 moles of $CaCl_2$ are dissolved to make 1500 mL of solution.

Answer Sheet
Calculate the molarity of the following solutions.

1) 2.2 moles of $LiNO_3$ are dissolved to make 750 mL of solution.

   2.9 M

2) 0.760 moles of $C_{12}H_{22}O_{11}$ are dissolved to make 2 L of solution.

   0.4 M

3) 4.4 moles of silver nitrate are dissolved to make 0.750 L of solution.

   5.9 M

4) 18 moles of $Cu(NO_3)_2$ are dissolved to make 1.250 mL of solution.

   14 M

5) 0.630 moles of Li_2SO_4 are dissolved to make 250 mL of solution.

   2.5 M

6) 6.1 moles of hydrogen chloride are dissolved to make 1.25 L of solution.

   4.9 M

7) 0.270 moles of potassium sulfate are dissolved to make 1.50 L of solution.

   0.180 M

8) 30 moles of sodium hydroxide is dissolved to make 1.750 mL of solution.

   20 M

9) 0.048 moles of sodium carbonate are dissolved to make 0.25 L of solution.

   0.19 M

10) 4.0 moles of $CaCl_2$ are dissolved to make 1500 mL of solution.

   2.7 M