Name_____

Calculate the molarity of the following solutions.

1) 494.7 grams of lithium sulfate are dissolved to make 1.50 L of solution.

2) 855.7 grams of sucrose are dissolved to make 750 mL of solution.

3) 342.3 grams of NH_4Cl are dissolved to make 1.75 L of solution.

4) 332.0 grams of sodium hydroxide are dissolved to make 1 L of solution.

5) 1,859.1 grams of KF is dissolved to make 2000 mL of solution.

6) 34.4 grams of $CaCl_2$ are dissolved to make 500 mL of solution.

7) 126.8 grams of iron(II) chloride are dissolved to make 0.250 L of solution.

8) 343.6 grams of ammonium sulfate are dissolved to make 1500 mL of solution.

9) 57.0 grams of NaCl are dissolved to make 1.5 L of solution.

10) 3,938.8 grams of $Cu(NO_3)_2$ are dissolved to make 1.250 mL of solution.

Answer Sheet

Calculate the molarity of the following solutions.

1) 494.7 grams of lithium sulfate are dissolved to make 1.50 L of solution.

3.00 M

2) 855.7 grams of sucrose are dissolved to make 750 mL of solution.

3.3 M

3) 342.3 grams of NH_4Cl are dissolved to make 1.75 L of solution.

3.66 M

4) 332.0 grams of sodium hydroxide are dissolved to make 1 L of solution.

8 M

5) 1,859.1 grams of KF is dissolved to make 2000 mL of solution.

20 M

6) 34.4 grams of $CaCl_2$ are dissolved to make 500 mL of solution.

0.6 M

7) 126.8 grams of iron(II) chloride are dissolved to make 0.250 L of solution.

4.00 M

8) 343.6 grams of ammonium sulfate are dissolved to make 1500 mL of solution.

1.7 M

9) 57.0 grams of NaCl are dissolved to make 1.5 L of solution.

0.65 M

10) 3,938.8 grams of $Cu(NO_3)_2$ are dissolved to make 1.250 mL of solution.

16.80 M