Name_____

Calculate the molarity of the following solutions.

1) 656.3 grams of hydrogen chloride are dissolved to make 2 L of solution.

2) 11 moles of copper(II) nitrate are dissolved to make 750 mL of solution.

3) 1,319.9 grams of NaOH is dissolved to make 1750 mL of solution.

4) 1,163.8 grams of $C_{12}H_{22}O_{11}$ is dissolved to make 1.50 L of solution.

5) 332.9 grams of calcium chloride are dissolved to make 1.25 L of solution.

6) 0.435 moles of $LiNO_3$ are dissolved to make 0.75 L of solution.

7) 0.350 moles of NH_4Cl are dissolved to make 1.75 L of solution.

8) 392.9 grams of iron(II) chloride are dissolved to make 1000 mL of solution.

9) 0.150 moles of Li_2SO_4 are dissolved to make 1 L of solution.

10) 52.3 grams of potassium sulfate are dissolved to make 2.0 L of solution.

Answer Sheet

Calculate the molarity of the following solutions.

1) 656.3 grams of hydrogen chloride are dissolved to make 2 L of solution.

9 M

2) 11 moles of copper(II) nitrate are dissolved to make 750 mL of solution.

15 M

3) 1,319.9 grams of NaOH is dissolved to make 1750 mL of solution.

18.9 M

4) 1,163.8 grams of $C_{12}H_{22}O_{11}$ is dissolved to make 1.50 L of solution.

2.27 M

5) 332.9 grams of calcium chloride are dissolved to make 1.25 L of solution.

2.40 M

6) 0.435 moles of $LiNO_3$ are dissolved to make 0.75 L of solution.

0.58 M

7) 0.350 moles of NH_4Cl are dissolved to make 1.75 L of solution.

0.200 M

8) 392.9 grams of iron(II) chloride are dissolved to make 1000 mL of solution.

3 M

9) 0.150 moles of Li_2SO_4 are dissolved to make 1 L of solution.

0.2 M

10) 52.3 grams of potassium sulfate are dissolved to make 2.0 L of solution.

0.15 M